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1. In covalent bonds, electron sharing usually occurs so that atoms attain the electron configuration of noble gases. (8) 2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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hyobinkim1310. Terms

in this set (26) a

depiction of the

arrangement of atoms

in molecules and

polyatomic ions.

structural formula. a

covalent bond in which

only one pair of

electrons is shared.

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bond. The chemical
bond that results from
sharing valence
electrons. Molecule.
Formed when two or
more atoms bond
covalently. Lewis
structure. Represent
the arrangement of
electrons in a
molecule. Sigma
bonds. Single covalent
bonds. Pi bond.

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molecular orbital that
lies along the axis
connecting two atom....
coordinate covalent
bond. a bond in which
one atom contributes
both bonding electrons
to.... bond dissociation
energy. the total
energy required to
break the bond
between two covale....
42 terms.

connorhumphrey2.

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Study ...

Technically, any covalent bond between two different elements is polar. However, the degree of polarity is important. A covalent bond between two different elements may be so slightly imbalanced that the bond is, essentially,

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nonpolar. A bond may be so polar that an electron actually transfers from one atom to another, forming a true ionic bond.

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Chapter 8 121 Section
8.1 The Covalent Bond
pages 240-247

Practice Problems page
244 Draw the Lewis
structure for each
molecule. 1. PH_3 H_2O H_2S
 $\text{H}-\text{H}$ $\text{H}-\text{H}$ P respectively,
for single, double, and
triple $\text{P}=\text{O}$ $\text{P}=\text{S}$ 2. H_2S H_2O
 $\text{H}-\text{H}$ $\text{H}-\text{S}-\text{S} \dots$

Covalent Bonding Covalent Bonding

A covalent bond in
which three pairs of

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electrons are shared by two atoms. Structural formula A chemical formula that shows the arrangement of atoms in a molecule or a polyatomic ion, each dash between a pair of atoms indicates a pair of shared electrons.

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Covalent Bonding

When two atoms of the same kind bond,

neither of them wants

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to lose or gain an electron. Therefore, they must share electrons. Each pair of shared electrons

constitutes one chemical bond.

Strengths of Covalent Bonds

Basic Concepts of Chemical Bonding

Covalent bonds form when electrons are shared between atoms and are attracted by the nuclei of both

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atoms. In pure covalent bonds, the electrons are shared equally. In polar covalent bonds, the electrons are shared unequally, as one atom exerts a stronger force of attraction on the electrons than the other.

Covalent Bonding - Chemistry

Valence bond theory describes a covalent bond as the overlap of

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half-filled atomic orbitals (each containing a single electron) that yield a pair of electrons shared between the two bonded atoms. We say that orbitals on two different atoms overlap when a portion of one orbital and a portion of a second orbital occupy the same region of space.

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Major topics: bond
energy, bond length,

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Chapter 8 (Bonding: General Concepts) - Part 1

8.0: Prelude to Covalent Bonding Yet oxygen demonstrates very different magnetic behavior than nitrogen. We can pour liquid nitrogen through a magnetic field with no visible interactions, while liquid oxygen is

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attracted to the magnet and floats in the magnetic field.

Answers 8: Advanced Theories of Covalent Bonding - Chemistry

...

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This video describes how atoms covalently bond, and form single, double or triple bonds. Pi bonds are discussed as well as bond strength. This video...

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A bond formed when two atoms share two electrons. A tightly bonded group of atoms that behaves as a unit and has a negative or positive charge. A neutral group of atoms joined together by covalent bonds. An atom where one side of the molecule is slightly

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negative and the other
is more positive.

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