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Chemistry 12 Worksheet 4 Ka

Chemistry 12 Worksheet 4-4—Ka & Kb Calculations

Chemistry 12 Worksheet 4-4 Ka and Kb Calculations Name _____

37 Due Date _____ Correct and Hand In by _____ NOTE: For this worksheet, you must show all of your steps in each calculation. State any assumptions clearly.

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Chemistry 12 Worksheet 4-4—Ka & Kb Calculations Chemistry

12-Worksheet 4-4—Ka & Kb Calculations Page 3 of 4 Pages 7.

Calculate the pH of a 0.22 M solution of the salt NaNO_2 . Show

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all of your steps clearly. (6 marks) Answer ____ 8. A 0.40 M solution of the lactate ion ($C_3H_5O_3^-$) (a weak base), has a pH of 8.728.

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Chemistry 12. Worksheet 4-4. Ka and Kb Calculations. Name ____ Due Date ____ Correct and Hand In by ____ NOTE: For this worksheet, you must show all of your steps in each calculation. State any assumptions clearly. Make sure your answer is in the correct number of significant digits as justified by the data and make sure your answer has the ...

Chemistry 12

Chemistry 12 Worksheet 4-4—Ka & Kb Calculations Chemistry 12-Worksheet 4-4—Ka & Kb Calculations Page 2 of 4 Pages 3. The pH in a 0.25 M solution of the acid HBrO is 4.65 . Using this, calculate the value of Ka for the acid HBrO. (4 marks) Answer

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_____ 4. The pH in a solution of benzoic acid is 2.355.

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Chemistry 12 Worksheet 4-4—Ka & Kb Calculations 9. The weak base ethylamine ($C_2H_5NH_2$) has a K_b of 6.4×10^{-4} a) Write the equilibrium equation for the ionization of ethylamine. (1 mark) b) What $[C_2H_5NH_2]$ is required to produce an ethylamine solution with a pH 12.102? pH: 11.102 / .B% (4 marks) +0.0126 S 0.0jÆ5 0.24"3 T 0.0065 Co Answer co

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Chemistry 12 Worksheet 4-4—Ka Kb Calculations 3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of K_a for the acid HBrO . (4 marks) 4.65 so $(-4.65) \cdot 10^{-5}$ w 4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks) $10^{-2.355} = 0.004416$ z

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Unit 4 notes part 3 Ka with the Unit 4 notes part 3 Ka include sample calc answers. There will be a QUIZ on Arrhenius Neutralization equations (worksheet 4.2), and two types of Ka Calc questions (worksheet 4.3) and Acid-Acid competition questions as they relate to Bronsted-Lowry Conjugate pairs (worksheet 4.4), on Friday Feb 28 (Term 3 mark).

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CHEMISTRY 12 WORKSHEET 4.10 Titrations and Titration Curves

1. Calculate the pH of the solution resulting from the following reactions: a) 30.0 ml of 0.20 M NaOH and 30.0 ml of 0.20 M HCl b) 30.0 ml of 0.20 M NaOH and 40.0 ml of 0.20 M HCl c) 50.0 ml of 0.10 M CH₃COOH and 50.0 ml of 0.10 M KOH d) 10.0 ml of 0.10 M HClO₄ and 10.0 ml of 0.10 M NH₃

2.

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Unit 5 Review - Study your notes and your Hebden textbook.

Unit 5 (Redox) - Chemistry 12

Determine the pH of a solution in which 1.00 mol H_2CO_3 ($K_a = 4.2 \times 10^{-7}$) and 1.00 mole NaHCO_3 are dissolved in enough water to form 1.00 L of solution. This is a buffer. The weak acid is carbonic acid and the conjugate base is bicarbonate

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Worksheet 2-3 - Calculations Involving the Equilibrium Constant

Page 8 has a 0.400. Exactly 1.00 mol of each gas was in a 100.0 L vessel and the K_c is 1.363. 3.0×10^{-2} mol of N_2 and 2.932×10^{-2} mol of O_2 are introduced into a 1.0 L container at 250°C, find K_c . The reaction. Chemistry 12 Unit 2 - Chemical Equilibrium

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