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ANSWER KEY 1 REVIEW QUESTIONS Chapter 8

Use only a periodic table to answer the following questions. 1.

Write complete electron configuration for each of the following elements: a) Aluminum (Al) $1s^2 2s^2 2p^6 3s^2 3p^1$ b) Sulfur (S) $1s^2 2s^2 2p^6 3s^2 3p^4$ c) Manganese (Mn) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$ 2.

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**Chemistry Chapter 8
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Answer Key

CHAPTER 8 REVIEW

Chemical Equations
and Reactions MIXED
REVIEW SHORT

ANSWER Answer the following questions in the space provided. 1. b A balanced chemical equation represents all the following except (a) experimentally established facts. (b) the mechanism by which reactants

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combine to form
products.

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Chemical Equations
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REVIEW SHORT
ANSWER Answer the
following questions in
the space provided. 1.
b A balanced chemical
equation represents all
the following except (a)

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experimentally established facts. (b) the mechanism by which reactants combine to form products.

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CHAPTER 8 REVIEW
Chemical Equations
and Reactions SECTION
3 SHORT ANSWER

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Answers

Answer the following questions in the space provided. 1. List four metals that will not replace hydrogen in an acid. Choose from Cu, Ag, Au, Pt, Sb, Bi, and Hg.

Modern Chemistry **Chapter 8 Review** **Answers Section 2**

CHAPTER 8 REVIEW
Chemical Equations
and Reactions MIXED
REVIEW SHORT

ANSWER Answer the

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following questions in the space provided. 1. b A balanced chemical equation represents all the following except (a) experimentally established facts. (b) the mechanism by which reactants combine to form products.

Modern Chemistry

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8113 The first step always is to draw a valid Lewis structure when predicting molecular structure. When resonance is possible, only one of the possible resonance structures is necessary to predict the correct structure because all resonance structures give the same structure. The Lewis

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Modern Chemistry
Chapter 8. chemical equation. skeleton equation. balanced equation. precipitate. a concise representation of a chemical reaction using chemical.... a chemical equation that does not indicate the relative amount.... an equation in which each side has the same number of atoms of....

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Chemistry Chapter 8.
covalent bond.
molecule. lewis
structure. sigma bond.
chemical bond that
results from sharing
valence electrons.
forms when 2 or more
atoms bond covalently.
representation of the
arrangements of
electrons in a
molecule. single
covalent bonds; occurs

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when the pair of
shared electron....

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prepare the heath
chemistry chapter 8
review answers to
retrieve all hours of
daylight is okay for
many people. However,
there are still many
people who also don't

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considering reading. This is a problem. But, past you can keep others to begin reading, it will be better. One of the books that can be

Heath Chemistry

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Answers

Introductory to Chemistry Chapter 8. The chemical composition of substance r.... The chemical composition

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Answers

of substance c.... What
are the four
observations that
ind.... How to write a
chemical equation.
Physical change.
Chemical change. Gas
is released ...
Precipitate - insoluble
solid is produced...

introductory chemistry chapter 8 Flashcards and Study Sets ...

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Matter and Change

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Answers

Vocab. diatomic. single covalent bond. double covalent bond. seven diatomic molecules (BrINClHOF) two atoms of the same element. 1 bond- 2e. 2 pairs of e- are shared. hydrogen, oxygen, nitrogen, flourine, chlorine, bromine, iodine. diatomic. two atoms of the same element. single covalent bond.

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Stoichiometry.
Stoichiometry. formula
weight. molecular
weight. percent
composition. The study
of quantitative
relations between
amounts of reactan....
The sum of the atomic
weights of the
elements that make up
a s.... The formula
weight of a substance

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Answers

that exists as individual
mo...

chemistry stoichiometry chapter 8 Flashcards and Study ...

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Chemical Equations
and Reactions SECTION
3 SHORT ANSWER

Answer the following questions in the space provided. 1. List four metals that will not replace hydrogen in an acid. Choose from Cu, Ag, Au, Pt, Sb, Bi, and Hg. 2. Consider the metals iron and silver, both listed in Table 3 on page 286 of the text.

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Which one

**8 Chemical
Equations and
Reactions**

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